

Carboxymethyl Cellulose-Graft-Polyacrylic Acid Hydrogel: A Cost Effective and Reusable Adsorbent for Cu²⁺

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A POROUS hydrogel composed of carboxymethyl cellulose grafted with polyacrylic acid CMC-g-PAA was prepared, by using methylenebisacrylamide (MBA) as crosslinking agent and ammonium persulfate (APS) as initiator. CMC-g-PAA hydrogel was evaluated as an adsorbent for removal of divalent copper ions from waste water. The effect of different adsorption parameters such as pH, contact time and initial copper ion concentration on removal process from aqueous solution was investigated. Adsorption isotherms for Cu²⁺ well fitted the Langmuir equation with maximum adsorption capacity 471 mg/g. The kinetics of adsorption process are appropriate to be described with pseudo second order model. The high adsorption capacity and kinetic results prove that CMC-g-PAA can be potentially used as a cost effective adsorbent to remove heavy metals from aqueous solutions.

Keywords: Carboxymethyl cellulose, Hydrogel, Cu²⁺ removal and Adsorption.

Heavy metal ions contaminating waste water is of great concern because of the harmful, acute and even fatal effect to the human beings^(1,2). Recent evidence shows that excessive consumption of copper can cause nausea, vomiting, respiratory difficulties, liver and kidney failure^(3,4). Conventional technologies for the removal of heavy metal ions from aqueous solutions like chemical precipitation, ion exchange, chemical oxidation/reduction, reverse osmosis and ultrafiltration have inherent disadvantages, such as less efficiency, high costs and further generation of toxic sludge or other waste products⁽⁵⁾. Therefore, searching for cheaper, higher efficient and more environment-friendly technologies are attracting considerable attention of scientists.

Adsorption process is emerging as an alternative to conventional methods for the effective removal of heavy metal ions and organic pollutants from effluent⁽⁶⁻⁸⁾. Recently, a great deal of attention has been directed towards the production of bioadsorbents from natural polymers, such as starch⁽⁹⁾, lignin⁽¹⁰⁾, cellulose^(1,8,11,12), chitin/chitosan^(13,14) and agricultural wastes⁽¹⁵⁾. These bioadsorbents are renewable and have many advantages over traditional

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adsorbents, such as cost-effective, biodegradability, non toxicity, environmentally friendly materials, hydrophilicity and good reusability⁽¹⁵⁻¹⁷⁾.

Hydrogels are cross-linked networks of hydrophilic polymers physically and/or chemically cross-linked to prevent their dissolution⁽¹⁸⁾. The porous structure for these materials, facility of the incorporation of different chelating groups, high internal surface area and high water content are extremely suitable properties to adsorb organic toxic materials and heavy metals⁽¹³⁾. Polysaccharides have been used for preparing a promising hydrogels for removal of metal ions^(16,17,19).

The present article represents an optimized synthesis of a superabsorbent hydrogel via crosslinking graft polymerization of acrylic acid onto carboxymethylcellulose. The grafting of acrylic acid, an anionic comonomer, will increase the anionic carboxyl group's density which is capable of forming complexes with Cu^{2+} ions. The adsorption isotherms were evaluated to elucidate the adsorption mechanism and adsorption capability of the hydrogel. Moreover, the adsorption capacity and the kinetic of copper adsorption on the CMC-g-PAA hydrogel were studied.

Experimental

Materials

CMC sodium salt (> 99.5%) with high viscosity was purchased from Fluka BioChemika. The viscosity of 4% CMC in water at 25 °C is 1000-1500 mPaS. Acrylic acid (99%) and N,N'-methylenebisacrylamide were purchased from Sigma Aldrich and used without further purification. The other chemicals, potassium persulphate and copper sulphate hexahydrate were analytical grade and used as received.

Preparation of CMC -g-PAA hydrogels

Graft copolymerization of AA onto CMC using MBA as chemical crosslinker, was carried out according to the following steps: 0.2 g CMC in 10ml double distilled water was mixed with different concentrations of AA and different concentrations of MBA. After bubbling N_2 gas for 20 min to deoxygenate the solution, APS as a redox initiator was added to the solution and then the mixture was heated until the gelation was completed. The resulting hydrogel was cut into disks and immersed in an excess amount of double distilled water for 3 days to remove the residual unreacted monomers. Swollen hydrogels were dried in a vacuum oven for 2 days at 30 °C to a constant weight.

Swelling studies

Swelling experiments of the CMC-g-PAA hydrogel were carried out in distilled water. A cubic sample of hydrogel with average size (2×2×2) mm was immersed in distilled water until equilibrium swelling was attained. The swollen weights of the hydrogel were determined, after removal of the surface liquid using tissue paper. The percent swelling was calculated by the following equation:

$$\% \text{Swelling} = 100[(W_t - W_0)/W_0] \quad (1)$$

where W_0 is the initial weight and W_t is final weight of the film at time t .

Adsorption experiments of Cu^{2+} ions

Stock solution of Cu^{2+} ions (1000 ppm) was prepared using $\text{CuSO}_4 \cdot 6\text{H}_2\text{O}$. The stock solution was then diluted to make solutions with desired concentrations. Batch adsorption experiments were conducted at room temperature in 100 ml bottles with continuous stirring by a magnetic stirrer. 50mg CMC-g-PAA were immersed in 100 ml solution with continuous stirring. After filtration, the concentration of Cu^{2+} remained in the solutions was determined by atomic absorption spectrophotometer PERKIN-ELMER 2380. Effect of pH on Cu^{2+} ion adsorption was investigated in the pH range 2-6 in solution with initial Cu^{2+} ions concentration of 100 ppm. The pH values selected in the experiments were chosen prior to the precipitation limit of the copper ions (pH = 6). The buffer solutions used for adjusting the pH of the medium were KCl/HCl (pH 2 and 3) and acetic acid/sodium acetate (pH 4, 5 and 6). The CMC-g-PAA gel was equilibrated at the certain pH for about 4 hr before the addition of Cu^{2+} ions. The amount of copper ion adsorbed on the hydrogel at adsorption equilibrium, q_e (mg/g), was calculated according to the following equation:

$$q_e = \frac{(C_0 - C_e)V}{w} \quad (2)$$

where C_0 and C_e are the initial and equilibrium metal ion concentrations (mg/L), V is the volume (L) of the metal ion solution used in the adsorption experiment, and w is the weight of the hydrogel disc (g). Data are representative of at least three experiments, and standard deviations are less than 6.0%.

Desorption and reusability behaviors of CMC-g-PAA hydrogel

A total of 50 mg of dry hydrogel discs were added to 100 ml of metal ion solution (200 mg/L), and the mixture was agitated continuously at the optimum pH and equilibrium time to reach its adsorption equilibrium. The hydrogel discs were then withdrawn from the solution and washed with deionized water 3 times. Thereafter, the hydrogel discs were immersed into HNO_3 solution (0.5 M) for 2hr under stirring to remove the adsorbed metal ions from the hydrogel and then treated with 0.01 M NaOH for 8 hr to completely convert COOH groups to COO^- groups. Finally, the hydrogel particles were thoroughly washed with deionized water and again used in the adsorption experiment. The adsorption-desorption process was repeated 5 times, and the amount of adsorbed and desorbed metal ions in every cycle was calculated according to Equation 2.

Characterization

Infrared spectroscopy

Both the CMC and CMC-g-PAA hydrogel before and after Cu^{2+} adsorption were characterized by FTIR (Mattson 5000 FTIR spectrometer) in the range of 600-4000 cm^{-1} using KBr discs.

X-ray diffraction

X-ray diffraction (XRD) patterns were recorded with an Empyrean Powder Diffractometer (CuK α , 0.154 nm) between 3 and 70° 2 θ with a step size of 0.01°/sec. Samples were mounted on a silicon support.

Scanning Electron Microscopy and Energy-Dispersive X-ray spectroscopy

SEM was carried out on a JEOL JXA-840A Electron probe microanalyzer with tungsten filament (30 kV). For EDX experiments an Oxford INCAx-sight SN detector with a resolution of 128 eV at 5.9 keV was used.

Results and Discussion

Optimization of the equilibrium swelling%

The data given in Fig.1A show that the highest equilibrium swelling value was achieved at 0.01 mol/L of MBA. However, with increasing MBA concentration, the equilibrium swelling gradually decreased, which may be due to increasing the crosslinking density. The high crosslinking density decreases space between co-polymer chains and consequently produces a rigid structure with no ability for expanding to its fullest extent and hence decreases the water absorbance.

The effect of monomer concentration on the equilibrium swelling of the resulting hydrogel was studied. As shown in Fig. 1B, increasing the content of AA led to gradual increasing in the swelling%, where it attained its maximum value (335%) at 1.7 mol/l then decreased. The swelling-loss after the maximum may be attributed to the higher concentrations of AA which could induce much more chain transfer and termination reactions leading to more homopolymerization instead of copolymerization. In addition, the increase in viscosity of the reaction medium may hinder the movement of the reactants and deactivate the macroradical growing chains after their formation⁽²⁰⁾.

Figure 1C illustrates the equilibrium swelling of CMC-g-PAA hydrogel as a function of APS concentration. Up to 0.0015 mol/l, equilibrium swelling increases gradually with APS increasing which may be attributed to increasing number of free radicals. These free radicals can lead to more chain ends on polysaccharide backbone and consequently produce higher graft polymerization extent which increases water absorbation. In contrast, subsequent decrease in water absorption after 0.0015 mol/L can be attributed to increasing the number of produced radicals which may enhance the terminating step via bimolecular collision. As there is an inverse relationship between Mw and initiator concentration, the increasing in initiator concentration will produce low molecular weight grafted PAA side chains in the formed hydrogel. These short PAA side chains have lower water absorbability. In addition, increasing the number of radicals enhances the crosslinking density described before as self-crosslinking phenomena⁽²¹⁾.

Figure 1D demonstrates the effect of the reaction temperature on the equilibrium swelling of CMC-g-PAA hydrogel. It has been observed that increasing the reaction temperature, beginning from 70°C and the water absorbency increased gradually with increasing temperature until 85 °C where it reached its maximum value (445%) and then decreased. The increase in swelling can be so explained that high temperatures favor the dissociation of the initiator leading to more copolymerization.

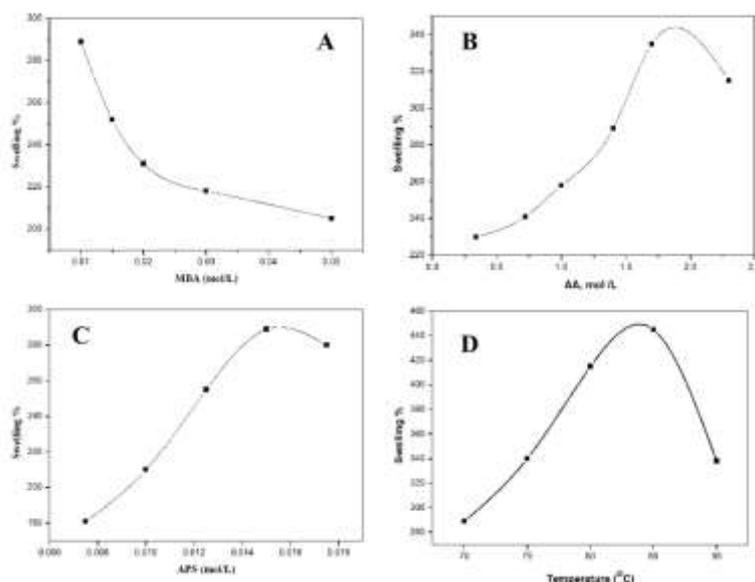


Fig.1. (A) Effect of crosslinker concentration on swelling % (AA 1.4 mol/L, APS 0.015 mol/L at 70 °C), (B) Effect of monomer concentration on swelling (MBA 0.01 mol/L, APS 0.015 mol/L at 70 °C), (C) Effect of initiator concentration on swelling % (MBA 0.01 mol/L, AA 1.4 mol/L at 70 °C), (D) Effect of temperature on swelling % (MBA 0.01 mol/L, AA 1.4 mol/L and APS 0.015 mol/L).

Cu²⁺ ion adsorption study

CMC-g-PAA hydrogel synthesized at 0.01 mol/L MBA, 1.7 mol/l AA, 0.0015mol/L APS and 85°C was evaluated as an adsorbent for Cu²⁺ ions in the following experiments. Experiments were conducted to determine the optimum conditions for the adsorption of Cu²⁺ ions on CMC-g-PAA hydrogel. It is expected that the network structure of CMC-g-AA hydrogel permits the adsorption and/or diffusion of the Cu²⁺. The effect of different parameters will be investigated.

Effect of pH of the initial Cu²⁺ solution

It is known that the pH is a critical parameter that can affect the hydrogel performance by influencing its swelling and ion uptake capability⁽⁸⁾. The effect of changing the pH of Cu²⁺ solution on adsorption capacity of the CMC-g-PAA

hydrogel is shown in Fig. 2. As can be seen the adsorption capacity increased sharply from 30 mg/g to a maximum adsorption of 75 mg/g with increasing pH of Cu^{2+} solution from 2 to 5 and then decreased with further increasing the pH to 6. This phenomenon can be explained as follows: at lower pH values the carboxylic groups are protonated and metal ions must compete with the protons. On the other hand at higher pH values, the carboxylic groups of the hydrogel are more ionized, therefore, formation of specific interaction with metal ions is more pronounced. Thus, copper ions are bounded to CMC-g-PAA hydrogel through an ion-exchange mechanism, which involves an electrostatic interaction between copper cations and negatively charged ligands. In addition, a more expanded polymeric network due to the repulsion between negatively charged COO^- groups at higher pH will also favor the diffusion and adsorption of copper ions in the interior of the hydrogel¹⁵. However, when the pH value is higher than 5 the adsorption capacity decreased, due to precipitation of heavy metal ions leading to reduction of its concentration in the aqueous solution⁽²²⁾.

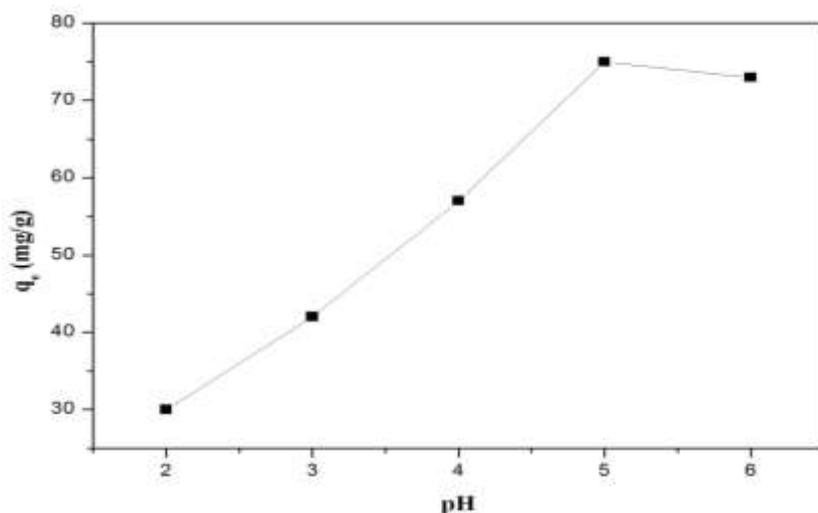


Fig. 2. Effect of pH on the adsorption of Cu^{2+} ions on CMC -g-PAA hydrogel (initial concentration 100 mg/L; CMC-g-AA 0.5 g/L; contact time 120 min; shaking rate 150 rpm).

Effect of contact time on the adsorption capacity

To establish the optimum contact time for maximum adsorption capacity and to determine the adsorption kinetics of CMC-g-PAA hydrogel, adsorption of copper ions capacity was measured as a function of time, and the results are shown in Fig. 3. The increase of the adsorption capacity of Cu^{2+} as contacting time progressed is remarkable and adsorption equilibrium could be achieved within 60 min. The short time period taken to reach the maximum adsorption equilibrium suggests an excellent affinity of the polymeric network toward Cu^{2+} ions in aqueous solutions. According to this result and the SEM morphologies in

Fig. 7, the following mechanism can be proposed to model the adsorption process: Rapid expansion of the polymeric network starts immediately after immersing the hydrogel in the buffer aqueous solution because of the highly hydrophilic network and strong electrostatic repulsion. Simultaneously, metal ions begin to transport toward the external surface of the hydrogel. This expansion may offer considerable channels for the diffusion of metal ions into the internal pores of hydrogel through the macropores, where they finally are adsorbed. The swelling of hydrogel which resulted from electrostatic repulsion between negatively charged carboxylate ions can favor the diffusion of Cu^{2+} ions and enhance the adsorption capacity of the adsorbent.

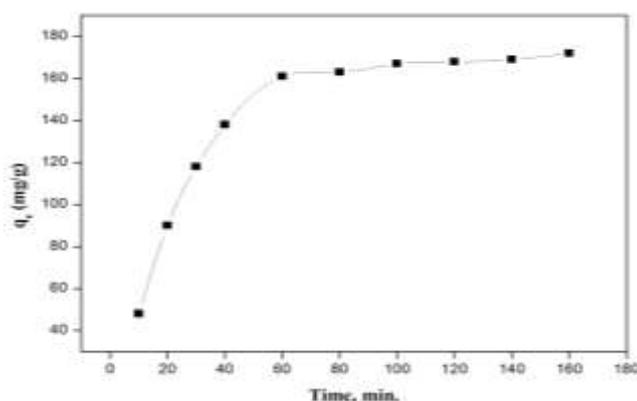


Fig. 3. Effect of contact time on the adsorption of Cu^{2+} ions on CMC-g-PAA hydrogel (initial concentration 100 mg/L; CMC-g-AA 0.5 g/L; pH 5; shaking rate 150 rpm).

To investigate the mechanism of the Cu^{2+} adsorption process, the two most commonly used kinetic models, *i.e.*, the pseudo-first-order and pseudo-second-order models were used to test the experimental data. The linear pseudo-first-order (Equation 3) and pseudo-second-order (Equation 4) are illustrated as follows:

$$\text{Log}(q_e - q_t) = \text{log}(q_e) - \frac{k_1}{2.303} t \quad (3)$$

$$\frac{t}{q_t} = \frac{t}{q_e} + \frac{k}{k_2 q_e^2} \quad (4)$$

where q_t (mg/g) is the adsorption capacity at time t (min), q_e (mg/g) is the adsorption capacity at equilibrium and k_1 (min^{-1}) and k_2 ($\text{g mg}^{-1} \text{min}^{-1}$) are the kinetics rate constants for the pseudo-first-order and pseudo-second-order models, respectively. Table 1 lists the computed results obtained from the first- and second-order kinetic models. Clearly, the correlation coefficient (R^2) for the pseudo-second-order kinetic model is higher than that of the pseudo-first-order kinetic model which suggests that this adsorption system is mainly a pseudo second-order kinetic adsorption process and the adsorption is a chemical process.

TABLE 1. Kinetic parameters for Cu²⁺ adsorption by CMC-g-PAA hydrogel.

Pseudo first order-model				Pseudo second- order model		
q _{e,exp} (mg/g)	q _{e,cal} (mg/g)	K ₁ (min ⁻¹)	R ²	q _{e,cal} (mg/g)	K ₂ (g mg ⁻¹ min ⁻¹)	R ²
171	152	0.0368	0.977	200	2.292×10 ⁻⁴	0.991

Effect of Cu²⁺ initial concentration and adsorption isotherms

The experimental data for the adsorption of Cu²⁺ on CMC-g-PAA hydrogel as a function of the initial Cu²⁺ concentration are shown in Fig.4. The results clearly indicate that with an increase in initial Cu²⁺ concentration (up to 400 mg/L), the adsorption capacity increased considerably, almost proportionally with the increment in the initial concentration of the metal ions. This may suggest that the adsorption sites on the gel were sufficient to entrap such amount of the transported metal from the bulk solution. However, at higher initial Cu²⁺ concentrations (>400 mg/l), the adsorption capacity no longer increased proportionally with the initial Cu²⁺ concentration, indicating that the number of the available ligands(carboxylic groups) in the superabsorbent hydrogel really became the limiting factor that controlled the amount of uptake ⁽²³⁾.

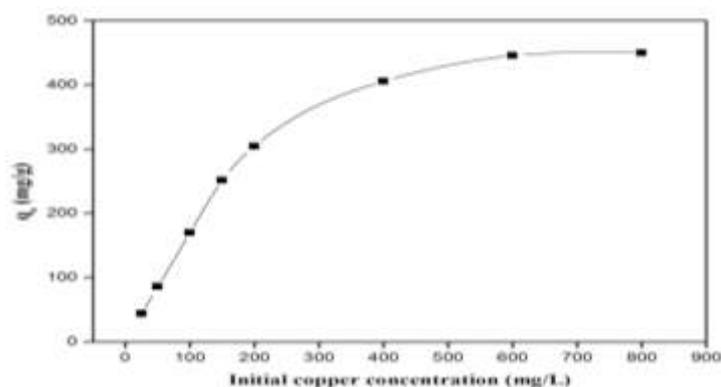


Fig. 4. Adsorption isotherm of Cu²⁺ by CMC -g-PAA hydrogel at various initial Cu²⁺ concentrations (initial concentration 25–800 mg/L, CMC-g-PAA 0.5 g/L, pH 5, shaking rate 150 rpm).

The adsorption isotherm was analyzed by fitting the equilibrium data with different isotherm models to investigate the mechanism of adsorption, to describe the interaction between adsorbate and adsorbent molecules and to estimate the extent of adsorption. In this study, Langmuir and Freundlich models were used to determine the adsorption equilibrium between the adsorbent and the metal ion. The isotherm constants for the two models were obtained by linear regression method. The isotherm data are presented in Table 2. The Langmuir ⁽²⁴⁾ isotherm model in the linearized form was calculated according to Equation 5:

$$\frac{C_e}{q_e} = \frac{K_s}{q_{\max}} + \frac{C_e}{q_{\max}} \quad (5)$$

where q_e is the equilibrium Cu adsorption amounts on the hydrogel (mg/g), C_e is the equilibrium Cu concentrations in the solution (mg/l), q_{\max} represents the maximum amount of Cu that could be adsorbed on the hydrogel (mg/g), and K_s is a constant of the Langmuir model (mg/l). The high correlation coefficient ($R^2 = 0.999$) of the linearized Langmuir equation indicates that the Langmuir model can describe the adsorption of Cu^{2+} on the CMC-g-PAA hydrogel.

TABLE 2. Parameters for Cu^{2+} ions adsorption by CMC-g-PAA hydrogel according to different equilibrium models.

Langmuir isotherm constants			Freundlich isotherm constants		
K_s (mg/L)	q_m (mg/g)	R^2	p (mg/g)	n	R^2
25.2	471	0.999	46.35	2.46	0.802

From the slope and the intercept of the straight line, the values of q_{\max} and K_s can be estimated to be 471 mg/g and 25.2 mg/l, respectively. The Langmuir isotherm model assumes that the adsorbed layer is one molecule in thickness and that all adsorption sites have equal energies and enthalpies of adsorption.

Freundlich model is an empirical model based on adsorption on a heterogeneous surface. The Freundlich equation is given by:

$$\log q_e = \frac{1}{n} \log C_e + \log P \quad (6)$$

where P is a constant representing the adsorption capacity (mg/g), and n is a constant depicting the adsorption intensity (dimensionless). The plot of $\log q_e$ versus $\log C_e$ for the experimental data illustrated that the correlation coefficient (R^2) for the Freundlich isotherm model is 0.802, which indicates that this model does not describe very well the adsorption processes of the composite hydrogel for Cu^{2+} . The values of the Freundlich model constants P and n are 46.35 and 2.46.

The essential features of Langmuir isotherm can be expressed in terms of a dimensionless separation constant factor, R_L , which is used to predict if an adsorption system is favorable or unfavorable. The R_L value can be estimated according to the following equation:

$$R_L = \frac{1}{1 + K_s C_0}$$

where K_s is the Langmuir equilibrium constant and C_0 (25-800 mg/L) is the initial concentration of Cu^{2+} ions. The calculated results revealed that the adsorption of Cu^{2+} ions on CMC-g-AA hydrogels is favorable because all R_L values were in the range of $0 < R_L < 1$.

Characterization of CMC-g-PAA before and after Cu²⁺ adsorption.

The structural changes of CMC and CMC-g-PAA hydrogel before and after copper adsorption were confirmed by FTIR spectroscopy (Fig. 5). Peaks at 1640, 2901 and 3438 cm⁻¹ are related to carboxyl group, asymmetric C-H stretching and O-H stretching vibration in CMC, respectively. However, CMC-g-PAA shows new peaks at 1728, 1550 and 1406 cm⁻¹ attributed to carbonyl stretching of the carboxylic acid groups and symmetric and asymmetric stretching modes of carboxylate groups, respectively. As the CMC-g-PAA hydrogel was loaded with Cu²⁺, the bond energy was altered by the chemical reaction in the adsorption process, which in turn resulted in the decrease of the spectral intensity for the 1710 stretching. This elucidated that Cu²⁺ had bound with the AA moieties of the CMC-g-PAA hydrogel⁽²⁵⁾. Moreover, the decrease in the intensity of the peak at 3438 cm⁻¹ after Cu²⁺ adsorption is an evidence for the binding of Cu²⁺ ions to the hydroxyl groups. These support that carboxyl groups of the AA moieties and the hydroxyl groups of the CMC are responsible for the Cu²⁺ adsorption on the hydrogel.

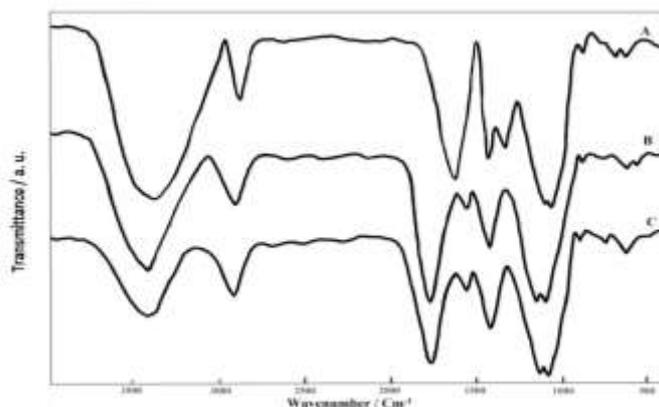


Fig. 5. FT-IR spectra of (A), CMC and CMC-g-PAA hydrogel (B), before and (C), after Cu²⁺ adsorption.

X-ray diffraction (XRD) is an efficient tool to investigate the degree of crystallinity for polymeric materials⁽²⁶⁾. XRD patterns for CMC and CMC-g-PAA before and after Cu²⁺ adsorption are represented in Fig. 6. CMC shows sharp reflections at 9.4 and 20.1° 2θ which are attributed to the partially crystalline structure of the neat CMC. However, these two bands became very marked broad in the CMC-g-PAA copolymer indicating its amorphous glassy structure. After Cu²⁺ adsorption, the broadening of the corresponding peaks was increased.

Figure 7 shows the morphological change of the hydrogel before and after the adsorption of Cu²⁺. CMC-g-PAA hydrogel exhibited a macroporous 3D structure with a pore diameter of about 80 μm (Fig. 7A). After the adsorption of Cu²⁺ ions, the expanded pores completely collapsed, and no pores could be observed on the surface of CMC-g-PAA hydrogel (Fig. 7B). The presence of COO⁻ groups in CMC-g-PAA hydrogel produces strong electrostatic repulsion in water, thus, the expansion of polymer chains occurs, forming a macroporous structure. The

interaction between electronegative COO^- groups and electropositive metal ions results in the electrostatic screening of anionic groups (COO^-) in the polymeric network, which leads to the decrease in electrostatic repulsion and the shrinkage of the network or pores. The expansion of the polymeric network and porous structure in the initial adsorption stage favors the diffusion and adsorption of metal ions from the exterior to the interior of the hydrogel. In addition, The results of EDX for the CMC-g-PAA before and after Cu^{2+} adsorption were displayed under SEM micrograph, it is difficult to extract clear evidence about Cu^{2+} in the adsorbed hydrogel, only atomic% for each sample is given.

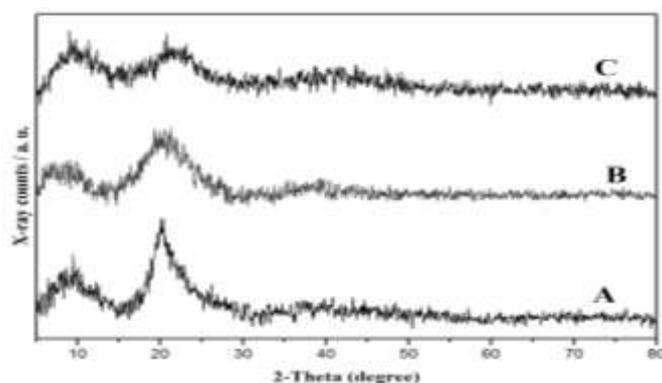


Fig. 6. XRD patterns of (A), CMC (B), CMC-g-PAA hydrogel and (C), copper adsorbed CMC-g-PAA hydrogel samples.

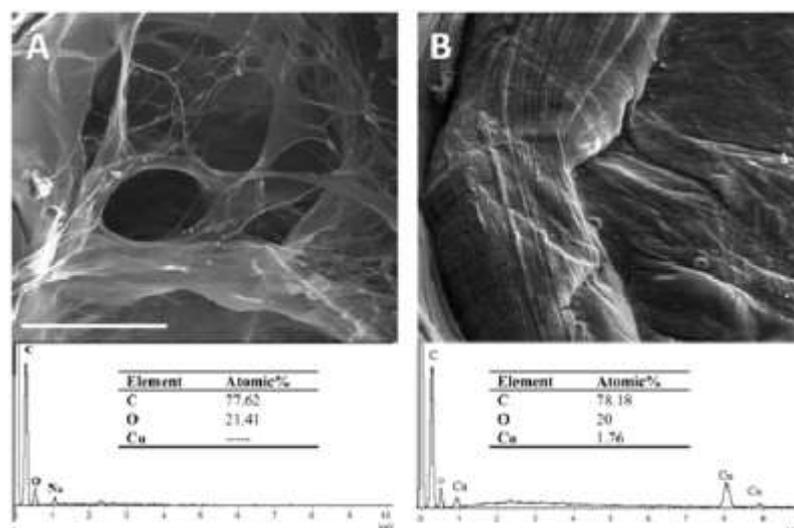


Fig. 7. SEM micrographs and corresponding representative EDX patterns of CMC-g-PAA hydrogel (A), before and (B), after the adsorption process in Cu^{2+} ions solutions (Scale bar = 200 μm).

Desorption and reusability

Table 3 illustrates the adsorption and desorption behaviors of CMC-g-PAA hydrogel for five cycles. The adsorption capacity of Cu^{2+} was 390 mg/g, after five cycles, indicating that reuse has little influence on the metal ion uptake. The recovery efficiency was high up to 98.4% in the first cycle and then slightly decreased to 96.3% after five cycles. In 0.5 M HNO_3 solution, the protons compete with metal ions for carboxyl groups, which is responsible for the easy desorption of metal ions. This again confirms that the main adsorption mechanism is ion exchange. Therefore, CMC-g-PAA hydrogel has stable physical and chemical properties and can be reused almost without significantly decreasing the adsorption capability after quite a number of cycles.

TABLE 3. The adsorption ratio of reused CMC-g-PAA hydrogel in consecutive five adsorption-desorption cycles.

	Adsorption capacity of Cu^{2+}	Recovery rate (desorption) of Cu^{2+}
Cycle 1	450	98.4
Cycle 2	443	98.2
Cycle 3	435	96.5
Cycle 4	420	96.3
Cycle 5	405	96.3
	390	

Conclusion

CMC-g-PAA hydrogel was synthesized and characterized as adsorbent for removal of Cu^{2+} from waste water. The experimental data showed that CMC-g-PAA prepared at optimum conditions exhibited high adsorption performance for copper ions in solution at pH 5. Equilibrium isotherm data were found to conform to Langmuir model with a good correlation. Kinetic study showed that the pseudo-second order model is appropriate to describe the adsorption process which demonstrates that the mechanism of adsorption includes mainly ionic interactions between Cu^{2+} and the carboxylate anions of the CMC-g-PAA hydrogel. The adsorption–desorption cycle results indicate that the regeneration and subsequent use would enhance the economics of the hydrogel applications in waste water treatment.

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هلاميات من الكربوكسى ميثيل سليولوز المطعم بحمض الأكريلك كمادة قليلة التكلفة ومتكررة الاستخدام لتنقية المياه من ايونات النحاس الثنائية

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فى هذه الدراسة تم تحضير هلاميات مسامية من الكربوكسى ميثيل سليولوز المطعم بحمض الاكريلك وذلك باستخدام الميثيلين بس اكريلاميد كرابط بين سلاسل البوليمرات والامونيوم بير سلفات كمادة بادئة للتفاعل. تم تقييم هذه الهلاميات كمادة مازة لايونات النحاس الثنائى من مياة الصرف . تم دراسة تأثير الأس الهيدروجينى وزمن التلامس وتركيز النحاس الثنائى على عملية الامتزاز. كما تم دراسة الايزوثيرم لعملية الامتزاز والذى اظهر ان القيمة العظمى لعملية الامتزاز قد تصل الى 471 مج/جرام . كما أظهرت الدراسة الكيناتيكية أن عملية الامتزاز من النوع ثنائى الدرجة الغير حقيقى. النتائج السابقة تثبت امكانية استخدام هلاميات الكربوكسى ميثيل سليولوز المطعم بحمض الأكريلك كمادة رخيصة للاستخدام المتكرر لازالة الايونات الثقيلة من مياة الصرف .